

1a i) $x = -0.2\text{m}$

1a ii) $m = 10^{-5}\text{g} = 10^{-8}\text{kg}$

$$t = 200 \text{ days} = 200 \times 24 \times 60^2 = 1.728 \times 10^7 \text{ s}$$

$$A = \left(\frac{0.05}{2} \times 10^{-3}\right)^2 \times \pi = 1.96 \times 10^{-9} \text{ kg m}^{-3}$$

$$M = \frac{10^{-8}}{1.96 \times 10^{-9}} = 5.102 \text{ kg m}^{-2}$$

$$c(x, t) = \frac{M}{\sqrt{4\pi Dmt}} e^{-\frac{(x-x_0)^2}{4Dmt}}$$

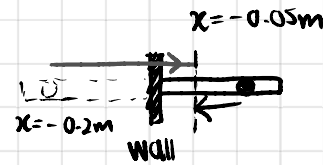
$$4Dmt = 4(10^{-9})(1.728 \times 10^7) \\ = 0.06912$$

$$c(-0.05, 1.728 \times 10^7) = c_{\text{real}} + c_{\text{image}}$$

$$= \frac{5.102}{\sqrt{0.06912 \times \pi}} e^{-\frac{(0.05-0)^2}{0.06912}} + \frac{5.102}{\sqrt{0.06912 \times \pi}} e^{-\frac{(0.15-0)^2}{0.06912}}$$

$$= (10.949 \times 0.96448) + (10.949 \times 0.72215)$$

$$= 10.56 + 7.9068 = 18.467 \approx 18.5 \text{ kg m}^{-3} \quad \#$$



10iii) Max concentration occurs at $x=0\text{m}$

$$C(0, 1.728 \times 10^{-7}) = 10.949$$

Assuming \rightarrow positive

$$J_{m'_{real}} = -D_m A \frac{dc}{dx}$$

$$\frac{A_{real}}{dx} = \frac{10.56 - 10.949}{-0.05 - 0} = 7.78$$

$$J_{m'_{real}} = -10^{-9} (1.96 \times 10^{-9}) (7.78) = -1.5249 \times 10^{-17}$$

Image

$$\frac{dc}{dx} = \frac{7.9068 - 10.949}{-0.05 - (-0.2)} = -20.281$$

$$J_{m'_{image}} = -10^{-9} (1.96 \times 10^{-9}) (-20.281) = 3.9751 \times 10^{-17}$$

$$\begin{aligned} \text{Overall flux: } & 10^{-17} (3.9751 + (-1.5249)) \\ & = 2.4502 \times 10^{-17} \approx 2.45 \times 10^{-17} \text{ kg s}^{-1} \end{aligned}$$

Direction: Mass flux is travelling from left to right.

2b i)

$$\begin{aligned}L &= L_0 e^{-\frac{k}{a}x} \\&= 500 e^{-0.2(5)} \\&= 183.94 \approx 184 \text{ mg/L.}\end{aligned}$$

(2b ii) $\frac{k_1}{k_2} = \beta^{(T_1 - T_2)}$

$$\frac{k_1}{0.2} = 1.065^{(25-20)}$$

$$k_1 = 0.27402$$

First 2 days 20°C , next 3 days 25°C .

$$L = L_0 e^{-k_2 \left(\frac{x}{a}\right)}$$

$$L = 500 e^{-0.2(2)} = 335.16$$

At start of day 3, $L = 335.16 \text{ mg/l}$.

$$L = 335.16 e^{-0.27402(3)} = 147.31$$

2ai)

discharge location ⇒	channel	$DO_{sat} = 8.4 \text{ mg/L}$	$u = 0.1 \text{ m/s}$
$DO = 1 \text{ mg/L}$	$DO = 6 \text{ mg/L}$	$k_1 = 0.15 \text{ /d}$	$k_2 = 0.174 \text{ /d}$
$BOD = 30 \text{ mg/L}$	$BOD = 20 \text{ mg/L}$	$= 1.7361 \times 10^{-6}$	$= 2.0139 \times 10^{-6}$
$Q = 0.05 \text{ m}^3/\text{s}$	$Q = 0.5 \text{ m}^3/\text{s}$		

$$x_c = \frac{u}{k_2 - k_1} \ln \left\{ \frac{k_2}{k_1} \left[1 - \left(\frac{k_2}{k_1} - 1 \right) \left(\frac{D_o}{L_o} \right) \right] \right\}$$

$$L_o = \frac{30(0.05) + 20(0.5)}{0.05 + 0.5} = 20.909$$

$$D_o = \frac{1(0.05) + 6(0.5)}{0.05 + 0.5} = 5.5455$$

$$D_o = DO_{sat} - D_o = 8.4 - 5.5455 = 2.8545$$

$$\frac{D_o}{L_o} = \frac{2.8545}{20.909} = 0.13652$$

$$\frac{u}{k_2 - k_1} = \frac{0.1}{10^{-6}(2.0139 - 1.7361)} = 359971$$

$$\frac{k_2}{k_1} = \frac{2.0139 \times 10^{-6}}{1.7361 \times 10^{-6}} = 1.16$$

$$x_c = 359971 \ln \left\{ 1.16 \left[1 - (1.16 - 1)(0.13652) \right] \right\}$$

$$= 359971 \ln(1.1347) = 45489$$

ans: 45489m

(2aii)

$$D = D_0 e^{-\frac{k_1}{u}x} + \frac{k_1 L_0}{k_2 - k_1} \left(e^{-\frac{k_1}{u}x} - e^{-\frac{k_2}{u}x} \right)$$

$$\frac{x}{u} = \frac{45489}{0.1} = 454890$$

$$e^{-k_1 \left(\frac{x}{u} \right)} = e^{-1.7361 \times 10^{-6} (454890)} = 0.45397$$

$$e^{-k_2 \left(\frac{x}{u} \right)} = e^{-2.0139 \times 10^{-6} (454890)} = 0.40008$$

$$\frac{k_1 L_0}{k_2 - k_1} = \frac{1.7361 \times 10^{-6} \times 20.909}{10^{-6} (2.0139 - 1.7361)} = 130.67$$

$$D = 2.8545 (0.45397) + 130.67 (0.45397 - 0.40008) \\ = 8.3377 \text{ (5sf)}$$

$$DO = DO_{\text{sat}} - D = 8.4 - 8.3377 = 0.0623 \text{ mg/L } \checkmark$$

$$(2aiii) \quad L = L_0 e^{-\frac{k}{u}x} = 20.909 e^{-1.7361 \times 10^{-6} (454890)} \\ = 9.492 \approx 9.49 \text{ mg/L } \checkmark$$

(2b i)

$$P^s = 1.25 \text{ g s}^{-1}$$

$$(Q^s \ll Q^o)$$

$$Q^o = 1.02 \text{ m}^3 \text{ s}^{-1} \rightarrow$$



$$P^o = 23 \text{ PPb}$$

$$= 23 \text{ Mg/L}$$

$$= 0.023 \text{ g m}^{-3}$$

$$SA = 8.9 \times 10^5 \text{ m}^2 \rightarrow \text{depth} = 9 \text{ m}$$

$$P = 13.2 \text{ PPb} = 13.2 \text{ Mg/L}$$

$$= 0.0132 \text{ g m}^{-3}$$

$$P = \frac{Q^o P^o + Q^s P^s}{Q^o + K_s V}$$

$$0.0132 = \frac{1.02(0.023) + 1.25}{1.02 + K_s(8.9 \times 10^5 \times 9)}$$

$$0.013464 + 105732 K_s = 0.02346 + 1.25$$

$$105732 K_s = 1.26$$

$$K_s = 1.1917 \times 10^{-5}$$

$$= 1.19 \times 10^{-5} \text{ s}^{-1} \quad \checkmark$$

(2b ii)

$$\text{fraction of P remaining in lake} = \frac{1}{1 + K_s \theta}$$

$$\theta = \frac{V}{Q^o} = \frac{8.9 \times 10^5 \times 9}{1.02} = 7852941$$

$$\frac{1}{1 + K_s \theta} = \frac{1}{1 + (7852941)(1.1917 \times 10^{-5})} = 0.010573$$

$$\text{fraction of P removed by lake} = 1 - 0.010573 = 0.98943$$

$$\approx 0.989 \quad \checkmark$$

(Q3a)

case: pulse input without chemical reaction

$$t = \frac{L}{v} = \frac{10(10^3)}{0.1} = 100\,000\text{ s} \\ \approx 1.1574 \text{ day} \approx 1.16 \text{ d } \checkmark$$

(Q3b)

$$\frac{c}{c_0} = e^{-\frac{t}{\bar{t}}}$$

$$\bar{t} = \frac{V}{Q} = \frac{40\,000}{0.1} = 400\,000\text{ s}$$

$$\frac{c}{c_0} = 0.1$$

$$0.1 = e^{-t/400\,000}$$

$$t = \ln(0.1) (-400\,000) \\ = 921034\text{ s} \approx 10.66 \text{ days} \approx 10.7 \text{ d } \checkmark$$

(Q3c)

In a non-ideal CSTR model, the pollutant molecules do not all have the same residence time due to reasons such as channeling of flow, short-circuiting, dead zones, restricted inlet or outlet effects, which happens in a naturally mixed lake.

$$(Q4a) \quad k_p = 8000 \times 1.75 = 14000 \text{ d}^{-1}$$

$$(Q4b) \quad \frac{1}{F_{\text{solid}}} = 1 + \frac{V_w}{V_{\text{solid}}} \frac{1}{k_p} + \frac{V_{\text{gas}}}{V_{\text{solid}}} \frac{H_{\text{cc}}}{k_p}$$
$$= 1 + \frac{5 \times 10^5}{45} \frac{1}{14000} + \frac{5 \times 10^9}{45} \frac{0.00275}{14000} = 23.619$$

$$F_{\text{solid}} = (23.619)^{-1} = 0.042339 \approx 0.0423$$

$$\frac{1}{F_{\text{water}}} = \frac{V_{\text{solid}} k_p}{V_w} + 1 + \frac{V_{\text{gas}} H_{\text{cc}}}{V_w} = \frac{45(14000)}{5 \times 10^5} + 1 + \frac{5 \times 10^9 \times 0.00275}{5 \times 10^5}$$
$$= 29.76$$

$$F_{\text{water}} = (29.76)^{-1} = 0.033602 \approx 0.0336$$

$$\frac{1}{F_{\text{gas}}} = \frac{V_{\text{solid}} k_p}{V_{\text{gas}} H_{\text{cc}}} + \frac{V_w}{V_{\text{gas}} H_{\text{cc}}} + 1 = \frac{45(14000)}{5 \times 10^9 \times 0.00275} + \frac{5 \times 10^5}{5 \times 10^9 \times 0.00275} + 1$$
$$= 1.0822$$

$$F_{\text{gas}} = (1.0822)^{-1} = 0.92404 \approx 0.924$$

$$F_{\text{solid}} = 0.0423, \quad F_{\text{water}} = 0.0336, \quad F_{\text{gas}} = 0.924$$

(Q4C)

$$F_{\text{fauna}} = \frac{V_{\text{fauna}} C_{\text{fauna}}}{V_{\text{solid}} C_{\text{solid}} + V_{\text{water}} C_{\text{water}} + V_{\text{air}} C_{\text{air}} + V_{\text{fauna}} C_{\text{fauna}}}$$

(5A)

1. Water softener
2. Copper pipe erosion

$$(5b\ i) \quad Sh = \frac{1.09}{\epsilon} Re^{0.333} Sc^{0.333}, \quad Sh = \frac{kd}{D}$$

$$Sc = \frac{\mu}{\rho D} = \frac{0.0131}{(1)(5 \times 10^{-6})} = 2620$$

$$v = 100 \text{ cm d}^{-1} \\ = 0.0011574 \text{ cm s}^{-1}$$

$$Re = \frac{\rho \epsilon v d}{\mu} = \frac{(1)(0.3)(1.1574 \times 10^{-3})(1)}{0.0131} = 0.026505$$

$$Sh = \frac{1.09}{0.3} (0.026505)^{0.333} (2620)^{0.333}$$

$$= 14.913$$

$$Sh = \frac{kd}{D}$$

$$k = \frac{14.913(5 \times 10^{-6})}{1} = 7.4565 \times 10^{-5} \text{ cm s}^{-1}$$

(shown,
k is between
 $6-8 \times 10^{-5} \text{ cm s}^{-1}$)

(5bii) specific surface area of the bed considering pore volume only

$$a = \frac{S}{\epsilon V}$$

160 naphthalene spheres per m^3
 $1 m^3 \Rightarrow 10^6 cm^3$

$$S = 160 \pi d^2 = 160 (\pi) (1)^2 = 502.65$$

$$\epsilon V = 0.3 \times 10^6 = 300000$$

$$a = \frac{502.65}{300000} = 0.0016755$$
$$= 0.00168 cm^2 / cm^3$$

(5biii)

$$\ln \left(\frac{C_s - C_x}{C_s - C_I} \right) = \frac{-k a x}{v}, \quad C_I = 0$$

$$\ln \left(\frac{C_s - C_x}{C_s} \right) = \frac{-7.4565 \times 10^{-5} \times 0.00168 \times 10 \times 100}{0.001574}$$

$$\frac{C_s - C_x}{C_s} = e^{-0.10823}$$

$$C_s - C_x = 0.89742 C_s$$

$$C_x = C_s - 0.89742 C_s$$

$$C_x = 0.10258 C_s \approx 0.105 C_s$$